CLASS : 11   **CHEMISTRY**   **MARKS :50**

I)ANSWER ALL THE QUESTION :   **5 X 2 = 10**

1. What do you understand by the term mole.

2. Define equivalent mass

3. State Pauli's exclusion principle?

4. How many orbitals are possible for n=4?

5. Define modern Periodic law.

II)ANSWER ALL THE QUESTION :   **5 X 3 =15**

6. What do you understand by the term oxidation number?

7. What is the empirical formula of the following?

i) Fructose ($C_6H_{12}O_6$) found in honey ii)Caffeine ($C_6H_{10}N_4O_2$) a substance found in tea and coffee.

8. Give the electronic configuration of Mn^{2+} and Cr^{3+}

9. what are isoelectronic ions? Give examples

10. How many radial and angular nodes for 2s, 4p, 5d and 4f orbitals exhibit

IV)ANSWER ALL THE QUESTION :   **5 X 5 =25**

11.Distinguish between oxidation and reduction?

12.What is the difference between molecular mass and molar mass?
    calculate the molecular mass and molar mass for Carbon monoxide.

13.Describe the Aufbau Principle.

14.Write short notes on the following.

15.Explain the pauli's method for the determination of ionic radius.